

TABLE II: SCALE UP OF PRECIPITATION REACTIONS

Run #	Chemical Formula of Reagent	Chemical Formula of Reagent	List the four ions present in reagents	Formula of Solid Product appearance in test tube	Formula of Aqueous Product appearance in test tube
				Soluble	Soluble
2	NaOH	FeCl <sub>2</sub>	$\text{Na}^{+}$ $\text{OH}^{-}$ $\text{Fe}^{+3}$ $\text{Cl}^{-}$	$\text{Fe}(\text{OH})_3$ very small orange	very cloudy orange
3	NaNO <sub>3</sub>	FeCl <sub>2</sub>	$\text{Na}^{+}$ $\text{NO}_3^-$ $\text{Fe}^{+2}$ $\text{Cl}^{-}$	solid white powder	yellow cloudy
5	Na <sub>2</sub> CO <sub>3</sub>	FeCl <sub>2</sub>	$\text{Na}^{+}$ $\text{CO}_3^{2-}$ $\text{Fe}^{+2}$ $\text{Cl}^{-}$	white orange powder	very cloudy orange
6	Na <sub>2</sub> PO <sub>4</sub>	FeCl <sub>2</sub>	$\text{Na}^{+}$ $\text{PO}_4^{3-}$ $\text{Fe}^{+2}$ $\text{Cl}^{-}$	yellowish white cloud	yellow cloudy
12	AgNO <sub>3</sub>	NaOH		very small cloud	cloudy brown
13	NaOH	CuSO <sub>4</sub>	N	very dark brown	very dark
17	AgNO <sub>3</sub>	Na <sub>2</sub> CO <sub>3</sub>		white powder	white cloudy
18	AgNO <sub>3</sub>	Na <sub>2</sub> PO <sub>4</sub>		yellow cloud	clear yellow
19	CuSO <sub>4</sub>	Na <sub>2</sub> CO <sub>3</sub>		small blue powder	cloudy blue
20	CuSO <sub>4</sub>	Na <sub>2</sub> PO <sub>4</sub>		very small blue particles	cloudy light blue

## Lab 9 2 Double Replacement Reaction

**Daniel Coit Gilman, Harry Thurston  
Peck, Frank Moore Colby**

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