

Chemistry 100 Final Study Guide

- **Chapter 1: Introduction**
 - a. Physical vs. Chemical properties.
 - 1) List 2 physical and 1 chemical property of magnesium.
 - b. Units of measurement.
 - 1) Can you distinguish between metric and U.S. units of measurement?
 - 2) What is 0.00000568 in scientific notation? What is 4.63×10^7 in regular notation?
 - c. Density
 - 1) What is the density of a 37g block of metal that takes up 15mL of volume?
- **Chapter 3: Atomic Structure**
 - a. Historical Perspective
 - 1) List the major scientific accomplishments of the following scientists:
 - (i) Michael Faraday
 - (ii) Joseph Thomson
 - (iii) Robert Millikan
 - (iv) Ernest Rutherford
 - (v) Neils Bohr
 - (vi) Erwin Schrodinger
 - b. Isotopes
 - 1) Define an isotope and give an example of several isotopes for hydrogen.
 - c. Structure of the atom
 - 1) Name the 3 main parts of an atom, and indicate where in the atom they'd be found.
 - 2) For the element sodium: how many energy levels? How many sublevels? What orbitals? How many electrons?
 - 3) Write the full electron configuration of silicon.
 - d. Electron configurations and trends on the Periodic Table
 - 1) Locate each of the following on the Periodic Table: alkali metals, alkaline earth metals, transition metals, halogens, noble gases, metals, non-metals.
- **Chapter 5: Chemical bonds**
 - a. Electron-dot structures
 - 1) Draw the electron-dot structure of an oxygen atom.
 - b. Octet Rule and Ions
 - 1) What charge of ions do magnesium, chlorine and nitrogen form?
 - c. Ionic compounds
 - 1) Name the following compounds: AlBr_3 , CaS , Na_3N
 - d. Covalent compounds
 - 1) Name the following compounds: SO_2 , PCl_5 , CH_4
 - e. Polyatomic molecules
 - 1) Name and write the formula for the 5 -ate polyatomic molecules we discussed.
 - f. Electron-dot formulas for molecules
 - 1) Write the electron-dot formula for each of the 6 -ate polyatomic ions we discussed.
 - g. VSEPR Model and Molecular shape
 - 1) Draw the electron-dot formula of ammonia (NH_3) and state its VSEPR shape.
 - h. Molecule polarity
 - 1) Which of the VSEPR molecule shapes are polar and which are non-polar?

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Rachel S Tattersall

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Course Success in the Undergraduate General Chemistry Lab

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