

Name _____

Molecular Geometries Lab Report Sheet

Formula	Lewis Structure	# of electron domains/Shape	Electron domain geometry	VSEPR	Hybridization of central atom	Polar?
CO ₂	$\text{O}=\text{C}=\text{O}$ 24 valence 2 charge clouds	0	linear 180°	linear	SP	NP
BF ₃	$\text{F}:\ddot{\text{B}}:\text{F}$ 24 valence 32 $\text{F}=\text{B}-\text{F}$	0	trigonal planar no		SP ₂	NP
SnCl ₂	$\text{Cl}=\text{Sn}-\text{Cl}:$ 16 valence 24	2	Bent <120		SP ₂	P
BF ₄ ⁻	$\text{F}:\ddot{\text{B}}:\text{F}$ $\text{F}=\text{B}-\text{F}$ F	0	tetrahedral 109.5		SP ₃	P
AsH ₃	$\text{H}-\text{As}-\text{H}$ 8 valence	4	Tetrahedral Pyramidal <109.5		SP ₃	P
H ₂ S	$\text{H}-\ddot{\text{S}}-\text{H}$	2	Bent <109.5		SP ₃	NP
NH ₃ ⁻	$[\text{H}-\ddot{\text{N}}-\text{H}]^-$	2	Bent <109.5		SP ₃	P
PCl ₅	$\text{F}:\ddot{\text{P}}:\text{F}$ $\text{F}=\text{P}-\text{Cl}:$ F	0	Tetrahedral Bipyramidal 120 + 90		SP ₃ d	P

Lewis Structure And Molecular Models Lab Answers

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Instructors Manual to Lab Manual Ralph Petrucci, William Harwood, Geoffrey Herring, 2001

Laboratory

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