

### Experiment 10: Electrodes with a Standard Solution

From: *Basic Laboratory Experiments*, Experiment 10, 45-50

The purpose of this experiment is to determine the standard potential for several cells consisting of  $\text{Zn} / \text{Zn}^{2+}$  and  $\text{Cu} / \text{Cu}^{2+}$  electrodes. Electrode potentials are a tool to support calculations such as those involving Gibbs free energy, Gibbs free energy of activation, and other thermodynamic quantities. The standard potential for a half-cell is defined as the potential of the half-cell when the concentration of the species in the half-cell is 1 M. The standard potential for a half-cell is defined as the potential of the half-cell when the concentration of the species in the half-cell is 1 M. The standard potential for a half-cell is defined as the potential of the half-cell when the concentration of the species in the half-cell is 1 M.

In this experiment, you will determine the standard potential for several electrodes and calculate the standard potential for the cell. The standard potential for a half-cell is defined as the potential of the half-cell when the concentration of the species in the half-cell is 1 M.

#### Objectives:

- to determine the standard potential for several electrodes and calculate the standard potential for the cell.
- to calculate the standard potential for the cell.
- to determine the standard potential for the cell.
- to determine the standard potential for the cell.
- to calculate the standard potential for the cell.

#### Materials:

- Standard solution
- Standard solution (1 M)
- Standard solution
- Standard solution
- Standard solution
- Standard solution
- Standard solution
- Standard solution
- Standard solution

#### Procedure:

- Standard solution
- Standard solution
- Standard solution
- Standard solution

#### Results:

- Standard solution
- Standard solution

#### Discussion:

- This is a good starting point.
- This is approximately 10 M. The standard potential for the cell is approximately 1.1 V. The standard potential for the cell is approximately 1.1 V. The standard potential for the cell is approximately 1.1 V.

# Heath Chemistry 11 Lab 6b Answers

**N Noddings**



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