

Chapter Assessment Chemistry: Matter and Change • Chapter 4

Reviewing Vocabulary

Match each definition in Column A with the term in Column B.

Column A

Column B

1. Radiation deflected toward the positively charged plate
2. Atoms with the same number of protons but different numbers of neutrons
3. High-energy radiation that has no charge and no mass
isotopes
4. The smallest particle of an element that retains the properties of that element
5. The weighted average mass of an element's isotopes
radiation
6. The center-most part of an atom where the protons and neutrons mass unit
are contained
radiation
7. Radiation deflected toward the negatively charged plate
atomic theory
8. The rays and particles emitted by radioactive material
from
9. Equal to 1.67×10^{-24} the mass of a carbon-12 atom
radioactivity
10. Rays of radiation traveling from the cathode to the anode
radioactive decay
11. Process by which unstable nuclei lose energy
12. States that all matter is composed of atoms
13. Process by which some substances spontaneously emit radiation
Complete and contrast each pair of related terms.
14. mass number, atomic number
15. nuclear reaction, nuclear equation

Understanding Main Ideas

Use the periodic table to identify each element described below.

1. atomic number 65
2. 70 protons
3. 46 protons and 44 electrons
4. atomic number 24
5. 21 protons
6. atomic number 53

In the spaces at the left, write true if the statement is true; if the statement is false, change the italicized terms to make it true.

7. An atomic nucleus contains its protons and electrons
8. Neutrons have no electrical charge
9. Some particles have a charge of 3^{+}
10. An alpha particle consists of two protons and two neutrons
Complete the table below.

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