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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction  $2\text{ag aq} \rightarrow 2\text{cr}_2\text{o}_7\text{ s}$  label 4 2 3

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web mar 11 2014 solution ionic equation  $\text{h aq} + \text{cl aq} + \text{na aq} + \text{hco}_3\text{ aq} + \text{na aq} + \text{cl aq} + \text{h}_2\text{o l} + \text{co}_2\text{ g}$  the spectator ions are cl and na net ionic equation  $\text{h aq} + \text{cl aq} + \text{na aq} + \text{hco}_3\text{ aq} + \text{na aq} + \text{cl aq} + \text{h}_2\text{o l} + \text{co}_2\text{ g} \rightarrow \text{h aq} + \text{hco}_3\text{ aq} + \text{h}_2\text{o l} + \text{co}_2\text{ g}$  here is a video on spectator ions

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web the six no 3 aq ions and the six na aq ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation  $3\text{ba}_2\text{ aq} + 2\text{po}_4\text{ aq} \rightarrow \text{ba}_3\text{po}_4\text{ s}$

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web net ionic equations an equation for a reaction in solution showing only those particles that are directly involved in the

chemical change spectator ions an ion that is not directly involved in a chemical reaction an ion that does not change oxidation number or composition during a reaction products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water h aq oh aq h 2 o l the spectator ions in koh aq and hno<sub>3</sub> aq are k aq and no<sub>3</sub> aq how to identify and list the

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web the ag aq and cl aq ions become agcl s but the na aq ions and the no<sub>3</sub> aq ions stay as na aq ions and no<sub>3</sub> aq ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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