

Study Guide – Unit 1 Periodic Table **ANSWERS**

Structure of an atom

- List the three subatomic particles which make up an atom. Include their location, charge, and atomic mass.
 - Electron Negative charge Electron Cloud
 - Proton Positive charge Nucleus
 - Neutron Neutral charge Nucleus
- In a neutral atom, which two subatomic particles are equal in number to each other?
Protons and Electrons
- What does the atomic number of an element indicate?
The number of protons in the atom of that element
- What is an isotope?
Atoms with the same number of protons, but different masses due to a different number of neutrons
- Will two different atoms of carbon always have the same number of protons?
Yes

Element key

13
Al
26.982
Aluminum

- What are the number of protons, neutrons, and electrons in an atom of aluminum?
 - Protons = 13
 - Neutrons = 14
 - Electrons = 13

Bohr Model

- How do you know how many energy levels (orbits) a Bohr model will have?
 - By the period the element is located in
- How many electrons can each energy level hold?
 $1^{\text{st}} - 2$ $2^{\text{nd}} - 8$ $3^{\text{rd}} - 18$

The Periodic Table

- How is the periodic table arranged? Increasing atomic number
- What are the vertical columns $\updownarrow$ called? Groups or Families
- What are the horizontal rows $\leftrightarrow$ called? Periods
- Why are elements put into groups? Same number of valence electrons
- What is a valence electron? Number of electrons on the outer energy level (shell)
- Know how many valence electrons each group has
- Which group of metals are highly reactive and why? Alkali Metals. Only 1 ve⁻ to lose
- Which group of non-metals are highly reactive and why? Halogens. Only 1 ve⁻ to gain
- Name some physical properties of metals: luster, ductile, malleable, conductor of heat and electricity
- Name some physical properties of non metals: no luster, brittle, not a good conductor
- What are metalloids and where are they found on the periodic table? Share properties of metals and metalloids. On zig-zag line starting at Boron,

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Table of Contents Excel Lesson 3 Study Guide With Answers

1. Understanding the eBook Excel Lesson 3 Study Guide With Answers
 - The Rise of Digital Reading Excel Lesson 3 Study Guide With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Excel Lesson 3 Study Guide With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Excel Lesson 3 Study Guide With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Excel Lesson 3 Study Guide With Answers
 - Personalized Recommendations

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- Excel Lesson 3 Study Guide With Answers and Bestseller Lists
- 5. Accessing Excel Lesson 3 Study Guide With Answers Free and Paid eBooks
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 - Excel Lesson 3 Study Guide With Answers Enhanced eBook Features
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 - Highlighting and Note-Taking Excel Lesson 3 Study Guide With Answers
 - Interactive Elements Excel Lesson 3 Study Guide With Answers
- 8. Staying Engaged with Excel Lesson 3 Study Guide With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Excel Lesson 3 Study Guide With Answers
- 9. Balancing eBooks and Physical Books Excel Lesson 3 Study Guide With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Excel Lesson 3 Study Guide With Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Excel Lesson 3 Study Guide With Answers
 - Setting Reading Goals Excel Lesson 3 Study Guide With Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Excel Lesson 3 Study Guide With Answers
 - Fact-Checking eBook Content of Excel Lesson 3 Study Guide With Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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