

Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. The charge on an ion is:
 - a. always positive.
 - b. always negative.
 - c. either positive or negative.
 - d. zero.

- _____ 2. According to the octet rule, a calcium atom has a tendency to:
 - a. lose one electron.
 - b. lose two electrons.
 - c. gain one electron.
 - d. gain two electrons.

- _____ 3. If a compound forms by ionic bonding, which is *not* true?
 - a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.
 - b. The net charge of the compound is zero.
 - c. The compound contains just two atoms, each of opposite charge.
 - d. Several ions group together in a tightly packed structure.

- _____ 4. The only property listed that is *not* characteristic of ionic compounds is:
 - a. high melting point.
 - b. hardness.
 - c. lack of crystal structure.
 - d. good conductor of electricity.

- _____ 5. Which formula listed below represents a polyatomic ion?
 - a. NH_4^+
 - b. H_2SO_4
 - c. NaCl
 - d. H_2O

- _____ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because:
 - a. ionic substances tend to vaporize at room temperature.
 - b. ionic substances are brittle.
 - c. attractive forces between ions are greater than the attractive forces between molecules.
 - d. the numbers of positive and negative charges are equal in an ionic compound.

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