

|  |  |  |   |
|--|--|--|---|
| $\text{C}_2\text{H}_4$ $8+4$<br># of electrons = <u>12</u><br>$\begin{array}{c} \text{H} & & \text{H} \\   & &   \\ \text{H} : \text{C} : & : & \text{C} : \text{H} \end{array}$   | $\text{C}_2\text{F}_4$ $8+28$<br># of electrons = <u>36</u><br>$\begin{array}{c} : \ddot{\text{F}} : & : \ddot{\text{F}} : \\   & &   \\ \text{C} : & : & \text{C} \\   & &   \\ : \ddot{\text{F}} : & : \ddot{\text{F}} : \end{array}$  | $\text{CO}$ $4+6$<br># of electrons = <u>10</u><br>$: \text{C} : : : \text{O} :$   | $\text{O}_2$ $6+6$<br># of electrons = <u>12</u><br>$\ddot{\text{O}} : : \ddot{\text{O}}$   |
| $\text{CO}_2$ $4+6+6$<br># of electrons = <u>16</u><br>$\ddot{\text{O}} : : \text{C} : : \ddot{\text{O}}$  | $\text{C}_2\text{H}_2$ (H C C H) $8+2$<br># of electrons = <u>10</u><br>$\text{H} : \text{C} : : \text{C} : \text{H}$  | $\text{N}_2$ $5+5$<br># of electrons = <u>10</u><br>$: \text{N} : : : \text{N} :$  | $\text{HCN}$ $1+4+5$<br># of electrons = <u>10</u><br>$\text{H} : \text{C} : : : \text{N} :$  |
| $\text{CN}^-$ $4+5+1$<br># of electrons = <u>10</u><br>$[: \text{C} : : : \text{N} :]^-$   | $\text{SO}_4^{2-}$ $6+24+2$<br># of electrons = <u>32</u><br>$\left[ \begin{array}{c} : \ddot{\text{O}} : \\ : \ddot{\text{O}} : \text{S} : \ddot{\text{O}} : \\ : \ddot{\text{O}} : \end{array} \right]^{2-}$   | $\text{PO}_4^{3-}$ $5+24+3$<br># of electrons = <u>32</u><br>$\left[ \begin{array}{c} : \ddot{\text{O}} : \\ : \ddot{\text{O}} : \text{P} : \ddot{\text{O}} : \\ : \ddot{\text{O}} : \end{array} \right]^{3-}$   | $\text{ClO}_3^-$ $7+18+1$<br># of electrons = <u>26</u><br>$\left[ \begin{array}{c} : \ddot{\text{O}} : \\ : \ddot{\text{O}} : \text{Cl} : \ddot{\text{O}} : \end{array} \right]^-$                                     |
| $\text{CO}_3^{2-}$ $4+18+2$<br># of electrons = <u>24</u><br>$\left[ \begin{array}{c} : \ddot{\text{O}} : \\ : \ddot{\text{O}} : \text{C} : \ddot{\text{O}} : \end{array} \right]^{2-}$ <p><i>Handwritten notes:</i> <math>\text{O}=\text{C}=\text{O}</math> with arrows pointing to the double bonds.</p> | $\text{NO}_3^-$ $5+18+1$<br># of electrons = <u>24</u><br>$\left[ \begin{array}{c} : \ddot{\text{O}} : \\ : \ddot{\text{O}} : \text{N} : \ddot{\text{O}} : \end{array} \right]^-$ <p><i>Handwritten notes:</i> <math>\text{O}=\text{N}=\text{O}</math> with arrows pointing to the double bonds.</p> | $\text{SO}_2$ $6+12$<br># of electrons = <u>18</u><br>$\begin{array}{c} : \ddot{\text{O}} : \text{S} : \ddot{\text{O}} : \\ \updownarrow \\ \text{O} : \text{S} : \text{O} : \end{array}$                        | $\text{O}_3$ (O O O)<br># of electrons = <u>18</u><br>$\begin{array}{c} \ddot{\text{O}} : : \ddot{\text{O}} : \ddot{\text{O}} : \\ \updownarrow \\ \ddot{\text{O}} : \ddot{\text{O}} : : \ddot{\text{O}} : \end{array}$ |
| $\text{SF}_6$ $6+42$<br># of electrons = <u>48</u><br>$\begin{array}{c} & \text{F} & & \text{F} & & \\ &   & &   & & \\ \text{F} & : & \text{S} & : & \text{F} & \\ &   & &   & & \\ & \text{F} & & \text{F} & & \end{array}$  | $\text{XeF}_4$ $8+28$<br># of electrons = <u>36</u><br>$\begin{array}{c} & \text{F} & & \\ &   & & \\ \text{F} & : & \text{Xe} & : & \text{F} & \\ &   & & \\ & \text{F} & & \end{array}$  | $\text{PCl}_5$ $5+35$<br># of electrons = <u>40</u><br>$\begin{array}{c} & \text{Cl} & & \\ &   & & \\ \text{Cl} & : & \text{P} & : & \text{Cl} & \\ &   & &   & & \\ & \text{Cl} & & \text{Cl} & & \end{array}$ | $\text{SeF}_4$ $6+28$<br># of electrons = <u>34</u><br>$\begin{array}{c} & \text{F} & & \\ &   & & \\ \text{F} & : & \text{Se} & : & \text{F} & \\ &   & & \\ & \text{F} & & \end{array}$                               |

# Lewis Dot Structures Worksheet Answer Key

**ML Yell**



## **Lewis Dot Structures Worksheet Answer Key:**

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