

## SYLLABUS :

- (i) Electrolytes and non-electrolytes.  
*Definitions and examples.*
- (ii) Substances containing molecules only, ions only, both molecules and ions.  
*Substances containing molecules only, ions only, both molecules and ions. Examples: relating their composition with their behaviour as electrolyte (strong and weak), non-electrolyte.*  
*Definition and explanation of electrolysis, electrolyte, electrode, anode, cathode, anion, cation, oxidation and reduction (on the basis of loss and gain of electrons).*
- (iii) An elementary study of the migration of ions, with reference to the factors influencing selective discharge of ions, illustrated by the electrolysis of: molten lead bromide; acidified water with platinum electrodes and aqueous copper (II) sulphate with copper electrodes; electron transfer at the electrodes.  
*The above electrolytic processes can be studied in terms of electrolyte used, electrodes used, ionization reaction, anode reaction, cathode reaction, use of selective discharge theory wherever applicable.*
- (iv) Application of electrolysis: electroplating with nickel and silver; purification of copper; choice of electrolyte for electroplating.  
*Reasons and conditions for electroplating, names of the electrolytes and the electrodes used should be given. Equations for the reactions at the electrodes should be given for electroplating, refining of copper.*
- (v) Acids, bases and salts as electrolytes; reference should be made to the activity series as indicating the tendency of metals, e.g., Na, Mg, Fe, Cu, to form ions.

## 6.1 INTRODUCTION

This chapter deals with a process called "Electrolysis" which establishes a relationship between electrical energy and chemical change.

The word electrolysis is made up of two words 'electro' meaning flow of electrons or electricity and 'lysis' meaning separating, i.e., bringing about a (chemical) change in a substance by passage of electricity.

Many substances allow electricity to pass through them and many do not. It has also been found that all the substances allowing passage of electricity do not undergo chemical change.

Most metallic elements like copper, aluminium, iron and all alloys, whether in solid state or in molten state, allow an electric current to pass through them without undergoing any chemical change. Non-metals (except graphite) like sulphur, phosphorus, etc., do not conduct electricity. Therefore, metals and alloys are called **conductors** and non-metals are called **non-conductors** (insulators) of electricity.

The conduction of a current by a conductor is due to the flow of electrons. Insulators do not possess mobile electrons.

Chemical compounds behave differently when an electric current is passed through them. Some compounds conduct electric current in aqueous state or in molten state but not in solid state and are decomposed by the passage of electric current. For example, the passage of electric current through copper chloride solution decomposes it to metallic copper and chlorine. Such chemical compounds which conduct electric current in molten or aqueous state are called **electrolytes**. Some compounds do not conduct electric current at all whether in solid state or in aqueous state. They are called **non-electrolytes**, e.g., sugar solution.

Chemical decomposition of a chemical compound can be brought about by using electric current. Such a chemical process is called **electrolysis**.

# Electrolysis Chapter For Class 10

**Fleeming Jenkin**



## **Electrolysis Chapter For Class 10:**

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