

## Additivity of Heats of Reaction: Hess's Law

### Advanced Study Questions:

Name \_\_\_\_\_

One of the reactions below is the same as the combination (addition) of the other two reactions. Therefore, according to Hess's Law, the heat of reaction,  $\Delta H$ , of that reaction should be equal to the sum of the heats of reaction for the other two.



Determine which of the three reactions above may result from the direct addition of the other two. Clearly show how two equations add up to give the third one. (*Use equations as they are: do not reverse, adjust coefficients, etc.*)

For the reaction obtained as a sum of two others (called the overall reaction), show how can its  $\Delta H$  be expressed using  $\Delta H$  for two other reactions? *You do not need values, use symbols only.*

In this experiment, sodium hydroxide is delivered in two ways, as solid or in an aqueous solution. Which sample contains **larger number of moles** of NaOH: 2.65 g solid NaOH or 55.2 mL of 1.15M NaOH solution? *Show full work including units.*

# Lab 35 Heats Of Reaction Answers

**G Thomas**



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