

Equation Writing and Predicting Products

Chemistry I Acc

Introduction:

If you examine your bicycle after it has been left out in the rain a number of times you will find that it has begun to rust. Rust is a slow chemical reaction of the iron in the bicycle with oxygen. If a piece of sodium is put in water, a much more rapid chemical reaction occurs. Sodium reacts with water to produce sodium hydroxide and hydrogen gas. During this reaction, enough heat is liberated to ignite the hydrogen gas, causing it to explode.

Chemists observe what is happening in a chemical reaction and try to describe it in language that is clear and simple. A chemical equation uses formulas and symbols to describe the substances involved in a reaction, the physical state of the substance, the use of a catalyst, and relative proportions. The general form of an equation is:



In this investigation, you will perform a series of reactions and make careful observations of the changes that occur. Using simple tests and your knowledge of chemistry, you will determine the identity of the products. With this information, you will write and balance chemical equations to describe the reactions.

Materials

goggles
magnesium ribbon
triangular file
copper strip
cobalt chloride paper

lab burner
watch glass
glycerin
bent glass tube
test tube tongs

crucible tongs
6 test tubes
wooden splints
rubber tubing
graduated cylinder

manganese(IV) oxide (MnO_2)
lead(II) nitrate ($\text{Pb}(\text{NO}_3)_2$) 0.1M
limewater ($\text{Ca}(\text{OH})_2$ solution)
ammonium carbonate ($(\text{NH}_4)_2\text{CO}_3$)

potassium iodide (KI) 0.1M
copper(II) carbonate (CuCO_3)
one-hole rubber stopper
hydrogen peroxide (H_2O_2) 10% solution

Safety Warnings

Wear your goggles at all times during this investigation. Avoid looking directly at the burning Magnesium. The bright light could seriously damage your eyes. Tie back loose hair. Hydrochloric acid is corrosive. Avoid direct contact with it. If you get some on your skin, immediately wash it off with copious amounts of water. Ammonia is a skin and respiratory irritant so avoid inhaling it deeply. Lead compounds are poisonous so be sure to avoid contact with skin. If contact occurs immediately wash it off with copious amounts of water and inform your teacher. Glass tubing breaks easily. Exercise caution when working with it and use glycerin.

Procedure:

1. Put on your goggles and wear them throughout the lab. For each of the reactions, record your observations in the data table. Observations should include any evidence that a chemical reaction has taken place; the results of any test performed on any gases produced; the appearance of the products; and any other relevant data.
2. Obtain a 2 cm long piece of Mg ribbon. This length is approximate. The exact size is not important. Holding the Mg ribbon with the crucible tongs, carefully light the ribbon with the lab burner. Once lit hold the Mg ribbon over a watch glass to collect any ash and debris. When the Mg is done burning turn off the burner. **CAUTION:** Make sure hair is tied back. The tongs will be hot even after the Mg has finished burning. Do not touch them for at least five minutes. Place them on the watch glass to cool. Do not look directly at the Mg while it burns.

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Cord



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Chemistry Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02

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Table of Contents Lab 25 Equation Writing And Predicting Products

1. Understanding the eBook Lab 25 Equation Writing And Predicting Products
 - The Rise of Digital Reading Lab 25 Equation Writing And Predicting Products
 - Advantages of eBooks Over Traditional Books
2. Identifying Lab 25 Equation Writing And Predicting Products
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lab 25 Equation Writing And Predicting Products
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lab 25 Equation Writing And Predicting Products
 - Personalized Recommendations
 - Lab 25 Equation Writing And Predicting Products User Reviews and Ratings
 - Lab 25 Equation Writing And Predicting Products and Bestseller Lists
5. Accessing Lab 25 Equation Writing And Predicting Products Free and Paid eBooks

- Lab 25 Equation Writing And Predicting Products Public Domain eBooks
- Lab 25 Equation Writing And Predicting Products eBook Subscription Services
- Lab 25 Equation Writing And Predicting Products Budget-Friendly Options
- 6. Navigating Lab 25 Equation Writing And Predicting Products eBook Formats
 - ePub, PDF, MOBI, and More
 - Lab 25 Equation Writing And Predicting Products Compatibility with Devices
 - Lab 25 Equation Writing And Predicting Products Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lab 25 Equation Writing And Predicting Products
 - Highlighting and Note-Taking Lab 25 Equation Writing And Predicting Products
 - Interactive Elements Lab 25 Equation Writing And Predicting Products
- 8. Staying Engaged with Lab 25 Equation Writing And Predicting Products
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lab 25 Equation Writing And Predicting Products
- 9. Balancing eBooks and Physical Books Lab 25 Equation Writing And Predicting Products
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lab 25 Equation Writing And Predicting Products
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Lab 25 Equation Writing And Predicting Products
 - Setting Reading Goals Lab 25 Equation Writing And Predicting Products
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Lab 25 Equation Writing And Predicting Products
 - Fact-Checking eBook Content of Lab 25 Equation Writing And Predicting Products
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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