

Equation Writing and Predicting Products

Chemistry I Acc

Introduction:

If you examine your bicycle after it has been left out in the rain a number of times you will find that it has begun to rust. Rust is a slow chemical reaction of the iron in the bicycle with oxygen. If a piece of sodium is put in water, a much more rapid chemical reaction occurs. Sodium reacts with water to produce sodium hydroxide and hydrogen gas. During this reaction, enough heat is liberated to ignite the hydrogen gas, causing it to explode.

Chemists observe what is happening in a chemical reaction and try to describe it in language that is clear and simple. A chemical equation uses formulas and symbols to describe the substances involved in a reaction, the physical state of the substance, the use of a catalyst, and relative proportions. The general form of an equation is:

Reactants -----> Products

In this investigation, you will perform a series of reactions and make careful observations of the changes that occur. Using simple tests and your knowledge of chemistry, you will determine the identity of the products. With this information, you will write and balance chemical equations to describe the reactions.

Materials

goggles	lab burner	crucible tongs
magnesium ribbon	watch glass	6 test tubes
triangular file	glycerin	wooden splints
copper strip	bent glass tube	rubber tubing
cobalt chloride paper	test tube tongs	graduated cylinder
manganese(IV) oxide (MnO_2)	potassium iodide (KI) 0.1M	
lead(II) nitrate ($Pb(NO_3)_2$) 0.1M	copper(II) carbonate ($CuCO_3$)	
limewater ($Ca(OH)_2$ solution)	one-hole rubber stopper	
ammonium carbonate ($(NH_4)_2CO_3$)	hydrogen peroxide (H_2O_2) 10% solution	

Safety Warnings

Wear your goggles at all times during this investigation. Avoid looking directly at the burning Magnesium. The bright light could seriously damage your eyes. Tie back loose hair. Hydrochloric acid is corrosive. Avoid direct contact with it. If you get some on your skin, immediately wash it off with copious amounts of water. Ammonia is a skin and respiratory irritant so avoid inhaling it deeply. Lead compounds are poisonous so be sure to avoid contact with skin. If contact occurs immediately wash it off with copious amounts of water and inform your teacher. Glass tubing breaks easily. Exercise caution when working with it and use glycerin.

Procedure:

1. Put on your goggles and wear them throughout the lab. For each of the reactions, record your observations in the data table. Observations should include any evidence that a chemical reaction has taken place; the results of any test performed on any gases produced; the appearance of the products; and any other relevant data.
2. Obtain a 2 cm long piece of Mg ribbon. This length is approximate. The exact size is not important. Holding the Mg ribbon with the crucible tongs, carefully light the ribbon with the lab burner. Once lit hold the Mg ribbon over a watch glass to collect any ash and debris. When the Mg is done burning turn off the burner. **CAUTION:** Make sure hair is tied back. The tongs will be hot even after the Mg has finished burning. Do not touch them for at least five minutes. Place them on the watch glass to cool. Do not look directly at the Mg while it burns.

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Cord

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Chemistry Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02

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Pattison, Susan Arena, Leo R. Best, 2014-01-15 The most comprehensive book available on the subject. Introduction to General Organic and Biochemistry 11th Edition continues its tradition of fostering the development of problem solving skills featuring numerous examples and coverage of current applications. Skillfully anticipating areas of difficulty and pacing the material accordingly this readable work provides clear and logical explanations of chemical concepts as well as the right mix of general chemistry, organic chemistry and biochemistry. An emphasis on real world topics lets readers clearly see how the chemistry will apply to their career.

Instructor's Resource Manual Silberberg, 1999-07 *Merrill Laboratory Chemistry* L.

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