

DEGRADATION OF ENERGY

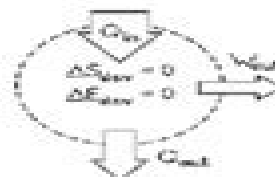
Questions and Problems

21.1. (a)



(b) Because the temperature of the gas does not change, neither does its thermal energy. Because the volume of the gas increases, its entropy increases. (c) The final state is different than the initial state, so it is not a steady device.

21.2.



21.3. In each case, some amount of mechanical energy is converted to thermal energy by friction, which increases the entropy of the snowboarder and her environment. Assuming that there is less friction on the ice than on the snow, the least amount of mechanical energy is converted to thermal energy in the third case, and the greatest amount of mechanical energy is converted to thermal energy in the second case, so the ranking is $\Delta S_3 < \Delta S_1 < \Delta S_2$.

21.4. There are two constraints on the operation of a steady device, because together with its environment, it must comply with both the energy law and the entropy law. To comply with the energy law, the sum of the energy inputs must equal the sum of the energy outputs. To comply with the entropy law, the entropy of the environment must not decrease. Thermal input of energy would cause a decrease in the entropy of the environment, and there is no thermal output of energy to compensate, so the we must have $Q_{out} = 0$, in which case the energy law gives $W_{out} = 320 \text{ J}$.

21.5. We know that the energy in must equal the energy out, so $Q_{out} + W_{out} = 425 \text{ J}$. Further, we know that Q_{out} cannot be exactly zero, since some heat energy must be delivered to the environment to allow the entropy law to be

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