

Chemistry - Chapter 9 Assessment Answers (Pages 281-282)

42. a) 2- b) 1+ c) 1- d) 3+

43. a) 2+ b) 2+ c) 3+ d) 1+

44. a) barium ion b) iodide ion c) silver ion d) mercury(II) ion

45. Cyanide,  $\text{CN}^-$  and hydroxide,  $\text{OH}^-$

46. a) hydroxide ion b) lead(IV) ion c) sulfate ion d) oxide ion

47. zero; A compound is electrically neutral.

48. The symbols for the cation and anion are written and the charges are balanced with subscripts. The name of the cation is followed by the name of the anion.

49. Determine the charge of the anion then work the formula backwards to find the charge of the transition metal cation needed to give a net charge of zero for the formula unit.

50. The symbols for the cation and anion are written and the charges are balanced with subscripts. Parentheses are used around the polyatomic ion if a subscript is needed. The name of the cation is followed by the name of the anion.

51. a and b

52. When more than a single polyatomic ion is needed to balance the formula.

53.  $\text{NH}_4\text{NO}_3$ , ammonium nitrate;  $(\text{NH}_4)_2\text{CO}_3$ , ammonium carbonate;  $\text{NH}_4\text{CN}$ , ammonium cyanide;  $(\text{NH}_4)_2\text{PO}_4$ , ammonium phosphate;  $\text{Sn}(\text{NO}_3)_2$ , tin(IV) nitrate;  $\text{Sn}(\text{CO}_3)_2$ , tin(IV) carbonate;  $\text{Sn}_2(\text{CN})_3$ , tin(IV) cyanide;  $\text{Sn}_2(\text{PO}_4)_3$ , tin(IV) phosphate;  $\text{Fe}(\text{NO}_3)_3$ , iron(III) nitrate;  $\text{Fe}_2(\text{CO}_3)_3$ , iron(III) carbonate;  $\text{Fe}(\text{CN})_3$ , iron(III) cyanide;  $\text{FePO}_4$ , iron(III) phosphate;  $\text{Mg}(\text{NO}_3)_2$ , magnesium nitrate;  $\text{MgCO}_3$ , magnesium carbonate;  $\text{Mg}(\text{CN})_2$ , magnesium cyanide;  $\text{Mg}_3(\text{PO}_4)_2$ , magnesium phosphate

54. two nonmetals

55. a) tri- b) mono- c) di- d) hexa- e) penta- f) tetra-

56. For formulas, write the correct symbols for each element with a subscript corresponding to the prefix before each element in the name. For naming, name each element in the order given. Use the subscript to determine the prefixes before each element in the name. The name ends in -ide.

57. a)  $\text{BCl}_3$  b) dinitrogen pentoxide c)  $\text{N}_3\text{H}_5$  d) carbon tetrachloride

58. a) hydrochloric acid b)  $\text{H}_2\text{SO}_4$  c) nitric acid d)  $\text{HC}_2\text{H}_5\text{O}_2$

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**John S. Phillips, McGraw-Hill Education**

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1. Understanding the eBook Glencoe Chemistry Chapter 9 Answers
  - The Rise of Digital Reading Glencoe Chemistry Chapter 9 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Glencoe Chemistry Chapter 9 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Glencoe Chemistry Chapter 9 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Glencoe Chemistry Chapter 9 Answers
  - Personalized Recommendations
  - Glencoe Chemistry Chapter 9 Answers User Reviews and Ratings
  - Glencoe Chemistry Chapter 9 Answers and Bestseller Lists
5. Accessing Glencoe Chemistry Chapter 9 Answers Free and Paid eBooks
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  - Glencoe Chemistry Chapter 9 Answers eBook Subscription Services
  - Glencoe Chemistry Chapter 9 Answers Budget-Friendly Options
6. Navigating Glencoe Chemistry Chapter 9 Answers eBook Formats

- ePub, PDF, MOBI, and More
- Glencoe Chemistry Chapter 9 Answers Compatibility with Devices
- Glencoe Chemistry Chapter 9 Answers Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Glencoe Chemistry Chapter 9 Answers
- Highlighting and Note-Taking Glencoe Chemistry Chapter 9 Answers
- Interactive Elements Glencoe Chemistry Chapter 9 Answers

8. Staying Engaged with Glencoe Chemistry Chapter 9 Answers

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- Participating in Virtual Book Clubs
- Following Authors and Publishers Glencoe Chemistry Chapter 9 Answers

9. Balancing eBooks and Physical Books Glencoe Chemistry Chapter 9 Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Glencoe Chemistry Chapter 9 Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Glencoe Chemistry Chapter 9 Answers

- Setting Reading Goals Glencoe Chemistry Chapter 9 Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Glencoe Chemistry Chapter 9 Answers

- Fact-Checking eBook Content of Glencoe Chemistry Chapter 9 Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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