

Practice Test: Gas Laws

1. Use Boyle's law to solve for the missing value in each of the following

a. $P_1 = 600\text{mm Hg}$ $V_1 = 200\text{mL}$ $P_2 = 780\text{mmHg}$ $V_2 = ?$

b. $P_1 = 2.4\text{ atm}$ $V_1 = 150\text{ mL}$ $P_2 = ?$ $V_2 = 435\text{mL}$

2. A sample of air has a volume of 550.0mL at 106°C. At what temperature will its volume be 700.0mL at constant pressure? Use Charles' law to solve this problem.

3. A sample of gas at 104°C and 0.870 atm occupies a volume of 3.0L. What volume would this gas occupy at 60°C and 1.7 atm?

4. A mixture of three gases A, B and C is at a total pressure of 8.15 atm. The partial pressure of gas A is 2.70 atm; that of gas B is 2.09atm. What is the partial pressure of gas C?

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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

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in Inorganic Chemistry Alexander Silverman, 1924 **Student Solutions Manual for Chang's Chemistry** Philip C. Keller, Raymond Chang, Jill L. Keller, 1988 National Note-book Sheets for Laboratory Work in Chemistry Arthur Stone

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