

## Chemistry EOC Review – Answer Key

1) a. 68,200,000 cg      b. 5000 nm      c. 0.0125 kL      d. 0.548 dm

2) ~21.5 mL

3) production of heat & light, formation of a gas, formation of a precipitate, color change

4)  $33.64 \text{ g} / 4.28 \text{ mL} = 7.86 \text{ g/mL} = \text{Fe}$

5) m.p.  $-183^{\circ}\text{C} = \text{CH}_4$  (methane)

6) H<sub>2</sub> – splint test (burning splint ignites hydrogen)  
O<sub>2</sub> – splint test (oxygen reignites glowing splint)  
H<sub>2</sub>O – cobalt chloride paper (paper changes from blue to pink in the presence of water vapor)  
CO<sub>2</sub> – splint test (carbon dioxide extinguishes burning splint);  
limewater test (limewater turns cloudy when carbon dioxide is bubbled through it, producing CaCO<sub>3</sub>)

7) The protons are **positively charged** particles found inside the **nucleus** along with the **neutrons**, which have no charge. The total number of these nucleons in an atom is the **mass # / atomic mass** of the atom. The electrons are the **negatively charged** particles found **orbiting in "clouds" around the nucleus**. Electrons can be gained by the atom to form **anions** or lost to form **cations**.

8) The thing that makes any element on the periodic table different from all other elements is the number of **protons** in its nucleus. **Isotopes** are atoms of the same type of element which have different **masses** because they have different numbers of **neutrons**.

9) a.  $^{39}_{19}\text{K}^{+}$   
b.  $^{57}_{26}\text{Fe}$   
c.  $^{32}_{16}\text{S}^{2-}$

10)

chemical symbol	atomic number	mass number	# of protons	# of electrons	# of neutrons	net charge
K <sup>+</sup>	19	42	19	18	23	+1
Sr	38	88	38	38	50	0
F <sup>-</sup>	9	20	9	10	11	-1
P	15	31	15	15	16	0
Ca <sup>2+</sup>	20	42	20	18	22	+2

11) uranium – 238      238 – 92 = 146 neutrons

12) (1) 199.59g

(2) They contain the same # of protons (they are the same element), but different #s of neutrons (electrons don't matter, b/c they have virtually no mass), & thus different masses.

# Eoc Chemistry Sample Items Answer Key

**Rachel Turoscy**

## **Eoc Chemistry Sample Items Answer Key:**

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*General Chemistry* Kenneth W. Whitten, Kenneth D. Gailey, Raymond E. Davis, 1988

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