

## Chapter 7: Ionic Compounds and Metals

### 7.1 Ion Formation p. 206

1. A chemical bond is the force that holds two atoms together.
2. Chemical bonds form by the attraction between the positive nucleus of one atom and the negative electrons of another atom.
3. Atom's try to form the octet—the stable arrangement of eight valence electrons in the outer energy level—by gaining or losing valence electrons.
4. A positively charged ion is called a cation.
5. An example is when sodium atom loses one valence electron to become a cation ion.
6. Metals are reactive because they lose valence electrons easily.
7. Transition metals commonly form 2+ or 3+ ions, but can form greater than 3+ ions.
8. Other relatively stable electron arrangements are referred to as pseudo-noble gas configurations.
9. An anion is a negatively charged ion.
10. An example would be chlorine gaining an electron to become a chlorine ion.
11. Nonmetal ions gain the number of electrons required to fill an octet.
12. Some nonmetals can gain or lose electrons to complete an octet.

### 7.2 Ionic Bonds and Ionic Compounds p.210

1. The electrostatic force that holds oppositely charged particles together in an ionic compound is called an ionic bond.
2. Compounds that contain ionic bonds are called ionic compounds.
3. Binary ionic compounds contain only two different elements—a metallic cation and a nonmetallic anion.
4. Formation of an ionic bond (table 7.4 p.211)
5. Positive and negative ions exist in a ratio determined by the number of electrons transferred from the metal atom to the non-metal atom.
6. The repeating pattern of particle packing in an ionic compound is called an ionic crystal.
7. The strong attractions among the positive and negative ions result in the formation of the crystal lattice.
8. A crystal lattice is the three-dimensional geometric arrangement of particles, and is responsible for the structure of many minerals.
9. Melting point, boiling point, and hardness depend on the strength of the attraction.
10. Melting and boiling points of some ionic compounds table 7.5 p.215.
11. In a solid, ions are locked into position and electrons cannot flow freely—solid ions are poor conductors of electricity.
12. Liquid ions or ions in aqueous solution have electrons that are free to move, so they conduct electricity easily.
13. An ion in aqueous solution that conducts electricity is an electrolyte.
14. Reactions that absorb energy are endothermic.
15. Reactions that release energy are exothermic.
16. The energy required to separate 1 mol of ions in an ionic compound is referred to as the lattice energy.
17. Lattice energy is directly related to the size of the ions that are bonded.

# Ionic Compounds And Metals Chapter 7 Study Guide Answers

**Susan McMurry**



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## **Table of Contents Ionic Compounds And Metals Chapter 7 Study Guide Answers**

1. Understanding the eBook Ionic Compounds And Metals Chapter 7 Study Guide Answers
  - The Rise of Digital Reading Ionic Compounds And Metals Chapter 7 Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Compounds And Metals Chapter 7 Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ionic Compounds And Metals Chapter 7 Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Compounds And Metals Chapter 7 Study Guide Answers
  - Personalized Recommendations
  - Ionic Compounds And Metals Chapter 7 Study Guide Answers User Reviews and Ratings
  - Ionic Compounds And Metals Chapter 7 Study Guide Answers and Bestseller Lists
5. Accessing Ionic Compounds And Metals Chapter 7 Study Guide Answers Free and Paid eBooks
  - Ionic Compounds And Metals Chapter 7 Study Guide Answers Public Domain eBooks
  - Ionic Compounds And Metals Chapter 7 Study Guide Answers eBook Subscription Services
  - Ionic Compounds And Metals Chapter 7 Study Guide Answers Budget-Friendly Options
6. Navigating Ionic Compounds And Metals Chapter 7 Study Guide Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Ionic Compounds And Metals Chapter 7 Study Guide Answers Compatibility with Devices
  - Ionic Compounds And Metals Chapter 7 Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ionic Compounds And Metals Chapter 7 Study Guide Answers
  - Highlighting and Note-Taking Ionic Compounds And Metals Chapter 7 Study Guide Answers
  - Interactive Elements Ionic Compounds And Metals Chapter 7 Study Guide Answers
8. Staying Engaged with Ionic Compounds And Metals Chapter 7 Study Guide Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Ionic Compounds And Metals Chapter 7 Study Guide Answers
- 9. Balancing eBooks and Physical Books Ionic Compounds And Metals Chapter 7 Study Guide Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ionic Compounds And Metals Chapter 7 Study Guide Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Ionic Compounds And Metals Chapter 7 Study Guide Answers
  - Setting Reading Goals Ionic Compounds And Metals Chapter 7 Study Guide Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ionic Compounds And Metals Chapter 7 Study Guide Answers
  - Fact-Checking eBook Content of Ionic Compounds And Metals Chapter 7 Study Guide Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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