

REAGENTS	NAME OF REACTION	WHAT IS PRODUCED	NOTES	KNOW THE MECH?
Mild oxidizing agents (see list)	Oxidation of Primary Alcohols (anhydrous)	$\text{R}-\text{CH}_2\text{OH} \longrightarrow \text{R}-\text{CHO}$ 1° alcohol → aldehyde	- Must use a mild oxidizing agent - Harsh oxidizing agents proceed to carboxylic acid	NO (Ch. 12.10)
1) O ₃ 2) DMS	Ozonolysis of Alkenes	$\text{R}-\text{CH}=\text{CH}-\text{R} \longrightarrow \text{R}-\text{CHO} + \text{R}-\text{CHO}$ disub alkene → aldehyde	Can also be used to form ketone when tetrasubstituted	NO (Ch. 8.12)
1) R ₂ BH 2) H ₂ O ₂ , NaOH	Hydroboration-oxidation of Terminal Alkyne	$\text{R}-\text{C}\equiv\text{CH} \longrightarrow \text{R}-\text{CH}_2\text{CHO}$ Terminal alkyne → aldehyde	Anti-Markovnikov addition of H₂O across π bond, then tautomerization	NO (Ch. 9.7)
Mild or Harsh oxidizing (see list)	Oxidation of Secondary Alcohols	$\text{R}-\text{CH}(\text{OH})-\text{R} \longrightarrow \text{R}-\text{C}(=\text{O})-\text{R}$ 2° alcohol → ketone	- Mild or Harsh oxidizing work. - Won't oxidize past ketone	NO (Ch. 12.1)

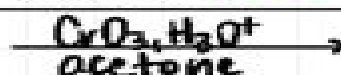
Oxidizing agents:

• 1° alcohol → aldehyde (mild stops) → carboxylic acid (harsh)
 • 2° alcohol → ketone (mild or harsh)

Mild:



Harsh:



$\text{H}_2\text{SO}_4, \text{H}_2\text{O}$ HgSO_4	Acid-catalyzed Hydration of Terminal Alkyne	$\text{R}-\text{C}\equiv\text{CH} \longrightarrow \text{R}-\text{C}(=\text{O})\text{CH}_3$ Terminal alkyne → ketone	Markovnikov addition of H₂O across π bond, then tautomerization	NO (Ch. 9.7)
$\text{Cl}-\text{C}(=\text{O})-\text{R}$ AlCl_3	Friedel-Crafts Acylation	$\text{C}_6\text{H}_6 \longrightarrow \text{C}_6\text{H}_5\text{C}(=\text{O})\text{R}$ Benzene → Benzoyl ketone	- Will not react if too deactivated. - Needs a Lewis acid present.	YES (Ch. 18.6)

Mastering Chemistry Ch 19 Key

John Olmsted, Gregory M. Williams



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