

CALCULATION PART

NAOH VS. OXALIC ACID

titration of 11th
chemistry

PRACTICAL

NaOH vs Oxalic Acid
Calculation part.

Given Data
Weight of



mole of

mole of

2 mole

molarity = $\frac{m}{V}$

$\frac{1}{2}$

= 1 mo

Lab 5titration Oxalic Acid

LP Steffe

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'LP Steffe' text.

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A Basic Laboratory Techniques

- 1 To cut a glass tube or glass rod
- 2 To bend the glass rod at an angle
- 3 To draw a glass jet from a glass tube
- 4 To bore a cork and fit a glass tube into it

B Characterisation and Purification of Chemical Substances

- 1 To determine the melting point of the given unknown organic compound and its identification simple laboratory technique
- 2 To determine the boiling point of a given liquid when available in small quantity simple laboratory method
- 3 To prepare crystals of pure potash alum $K_2SO_4 \cdot Al_2(SO_4)_3 \cdot 24H_2O$ from the given impure sample
- 4 To prepare the pure crystals of copper sulphate from the given crude sample
- 5 To prepare pure crystals of benzoic acid from a given impure sample

C Measurement of pH Values

- 1 To determine the pH value of vegetable juices fruit juices tap water and washing soda by using universal pH paper
- 2 To determine and compare the pH values of solutions of strong acid HCl and weak acid CH_3COOH of same concentration
- 3 To study the pH change in the titration of strong base Vs strong acid by using universal indicator paper
- 4 To study the pH change by common ion CH_3COO^- ion in case of weak acid CH_3COOH
- 5 To determine the change in pH value of weak base NH_4OH in presence of a common ion NH_4^+

D Chemical Equilibrium

- 1 To study the shift in equilibrium between ferric ions and thiocyanate ions by changing the concentrations of either of the ions
- 2 To study the shift in equilibrium between Co^{2+} and Cl^- ions by changing the concentrations of either of the ions

E Quantitative Analysis

- 1 To prepare M 10 oxalic acid solution by direct weighing method
- 2 To prepare M 10 solution of sodium carbonate by direct weighing method
- 3 To determine the strength of given solution of sodium hydroxide by titrating it against N 10 or M 20 solution of oxalic acid
- 4 To determine the strength of a given solution of hydrochloric acid by titrating it against a standard N 10 or M 20 sodium carbonate solution

F Qualitative Analysis

- 1 Analysis of Anions
- 2 Analysis of Cations

G Detection of Elements in Organic Compounds

- 1 To detect the presence of nitrogen sulphur and halogens in a given organic compound by Lassaigne's test
- 2 To detect the presence of nitrogen sulphur and halogens in the given organic compound sample number by Lassaigne's test

INVESTIGATORY PROJECTS

A Checking of Bacterial Contamination in Water

- 1 To check the bacterial contamination in drinking water by testing sulphide ions

B Methods of Water Purification

- 1 To purify water from suspended impurities by using sedimentation
- 2 To purify water by boiling
- 3 To purify water by distillation method
- 4 To purify water by

reverse osmosis technique 5 To purify water by GAC method 6 To purify water by bleach treatment 7 To purify water by oxidising agent 8 To purify water by ozone treatment method C Water Analysis 1 To test the hardness of different water samples D Foaming Capacity of Various Soaps 1 To compare the foaming capacity of different washing soaps 2 To study the effect of addition of sodium carbonate on foaming capacity of washing soap E Tea Analysis 1 To study the acidity of different samples of tea leaves tea by using pH paper F Analysis of Fruits and Vegetable Juices 1 To analyse the fruit and vegetable juices for the constituent present in them G Rate of Evaporation 1 To study the rate of evaporation of different liquids H Effect of Acids and Bases on Tensile Strength of Fibres 1 To compare the tensile strength of natural fibres and synthetic fibres 2 To study the effect of acids and bases on tensile strength of different fibres Log Antilog Table Hygienic Laboratory bulletin. no. 95-100, 1914 ,1914 **Laboratory Manual for the Course in Beginning Quantitative Analysis** Hobart Hurd Willard,1922 A Manual of Clinical Diagnosis by Means of Laboratory Methods Charles Edmund Simon,1922 Indexes to the Oak Ridge National Laboratory Master Analytical Manual Oak Ridge National Laboratory,1966 *The Chemical News and Journal of Physical Science* ,1912 **Chemical News and Journal of Physical Science** ,1895 *Bulletin of the Johns Hopkins Hospital* ,1915 *Experiment station record* ,1896 Laboratory Manual of Elementary Colloid Chemistry Emil Hatschek,1926 **Chemical News and Journal of Industrial Science** ,1903 *Advanced Research on Material Science, Environmental Science and Computer Science* Helen Zhang,David Jin,2011-08-16 Selected peer reviewed papers from the 2011 International Conference on Material Science Environmental Science and Computer Science MSESCS 2011 September 24 25 2011 Guangzhou China **The Pearson Guide to Objective Chemistry for the AIEEE** Singhal Atul,2009 To clear the All India Engineering Entrance Examination AIEEE students need to have a solid conceptual framework as well as adequate experience in solving original exam like questions The Pearson Guide to Objective Chemistry for the AIEEE seeks to serve this purpose by striking a unique balance between theory and practice Features such as Facts to Remember Important Guidelines Tools and Summary furnish the theoretical basis whereas practice questions arranged in levels sharpen the student s problem solving skills Designed and chiseled specifically for the AIEEE this book is the most focused manual for aspirants available Johns Hopkins Hospital Bulletin Johns Hopkins Hospital,1914 **The Journal of Biological Chemistry** ,1965 Vols 3 140 include the society s Proceedings 1907 41

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Table of Contents **Lab 5titration Oxalic Acid**

1. Understanding the eBook **Lab 5titration Oxalic Acid**
 - The Rise of Digital Reading **Lab 5titration Oxalic Acid**
 - Advantages of eBooks Over Traditional Books
2. Identifying **Lab 5titration Oxalic Acid**
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an **Lab 5titration Oxalic Acid**
 - User-Friendly Interface
4. Exploring eBook Recommendations from **Lab 5titration Oxalic Acid**
 - Personalized Recommendations
 - **Lab 5titration Oxalic Acid** User Reviews and Ratings
 - **Lab 5titration Oxalic Acid** and Bestseller Lists

5. Accessing Lab 5titration Oxalic Acid Free and Paid eBooks
 - Lab 5titration Oxalic Acid Public Domain eBooks
 - Lab 5titration Oxalic Acid eBook Subscription Services
 - Lab 5titration Oxalic Acid Budget-Friendly Options
6. Navigating Lab 5titration Oxalic Acid eBook Formats
 - ePub, PDF, MOBI, and More
 - Lab 5titration Oxalic Acid Compatibility with Devices
 - Lab 5titration Oxalic Acid Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lab 5titration Oxalic Acid
 - Highlighting and Note-Taking Lab 5titration Oxalic Acid
 - Interactive Elements Lab 5titration Oxalic Acid
8. Staying Engaged with Lab 5titration Oxalic Acid
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lab 5titration Oxalic Acid
9. Balancing eBooks and Physical Books Lab 5titration Oxalic Acid
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lab 5titration Oxalic Acid
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Lab 5titration Oxalic Acid
 - Setting Reading Goals Lab 5titration Oxalic Acid
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Lab 5titration Oxalic Acid
 - Fact-Checking eBook Content of Lab 5titration Oxalic Acid
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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