

## Gas Law Formula

| Gas Law  | Formula   | Description   |
|--|---|---|
| Boyle's Law  | $P_1V_1 = P_2V_2$                                       | At constant $T$ , as pressure increases, volume decreases.            |
| Charles' Law   | $\frac{V_1}{T_1} = \frac{V_2}{T_2}$                     | At constant $P$ , as volume increases, temperature increases.         |
| Gay-Lussac's Law   | $\frac{P_1}{T_1} = \frac{P_2}{T_2}$                     | At constant $V$ , as pressure increases, temperature increases.       |
| Combined Law   | $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$               | Obtained by combining Boyle's Law, Charles' Law and Gay-Lussac's Law. |
| Ideal Gas Law  | $PV = nRT$  |   |
| $V$ = volume in $\text{dm}^3$<br>$T$ = temperature in $\text{K}$ | $P$ = pressure in $\text{kPa}$<br>$n$ = number of moles | $R$ = ideal gas constant  |

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**William Simon,Daniel Base**

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