

Gas Law Formula

| Gas Law | Formula | Description |
|---|---|--|
| Boyle's Law | $P_1V_1 = P_2V_2$ | At constant T , as pressure increases, volume decreases. |
| Charles' Law | $\frac{V_1}{T_1} = \frac{V_2}{T_2}$ | At constant P , as volume increases, temperature increases. |
| Gay-Lussac's Law | $\frac{P_1}{T_1} = \frac{P_2}{T_2}$ | At constant V , as pressure increases, temperature increases. |
| Combined Law | $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$ | Obtained by combining Boyle's Law, Charles' Law and Gay-Lussac's Law. |
| Ideal Gas Law | $PV = nRT$ | |
| V = volume in dm^3 T = temperature in K | | P = pressure in kPa n = number of moles R = ideal gas constant |

Gas Laws Guide

William Simon, Daniel Base



Gas Laws Guide:

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