

Gas Laws Study Guide

Test Format:

- Multiple choice and problems

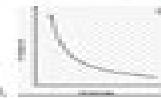
Items that will be provided for the test (read important notes at the bottom):

- Gas law equations
- The value of R
- Pressure conversion factors
- I will provide the equations on the test. You are responsible for knowing which laws the equations describe (ex: I will give you $P_1 \times V_1 = P_2 \times V_2$ and you must know that this is Boyle's Law).

- What is the name of the model we currently use to explain the behavior of gases? kinetic molecular theory (KMT).
- Define the kinetic molecular theory (KMT). **Gases are in constant motion with elastic collisions.**
- The KMT makes many assumptions about the **size**, **motion**, and **energy** of gas particles.
- According to the KMT, do gas particles experience attractive or repulsive forces? Explain why or why not. **Neither- they are too far apart**
- Completely describe the motion of gas particles according to the KMT. **Constant, Random, Straight paths, different velocity**
- Define elastic collision. **No kinetic energy is lost- just transferred between particles**
- What determines the kinetic energy of a particle? **Mass and velocity**
- In a sample of a single gas, do all particles have the same kinetic energy? Explain why or why not. **No – they have a different velocity**
- Define temperature. **Measurement of kinetic energy**
- Explain why gases have a low density. **Lots of space between particles (large volume)**
- Why are gases compressible? **Lots of empty space between particles**
- Why do gases expand to fill a container? **No attractive or repulsive forces**
- Why can gas particles flow easily past each other? **No attractive forces**
- Define diffusion. **Movement of one material through another**
- Define effusion. **Gas escapes through tiny openings**
- Describe how diffusion depends on the mass of the particles. **Diffusion is dependent on mass, so less mass more diffusion**
- Use the KMT to explain why a balloon expands as you blow air into it. **Particles are in-constant motion and collide with their surroundings expanding the balloon**

Boyle's Law: Answer the following questions about Boyle's Law.

- Describe Boyle's Law in your own words. **As pressure increases, volume decreases**
- What is the formula? $P_1V_1 = P_2V_2$
- What are the variables? What are the possible units used in the formula? **Pressure and volume, atm, kPa, mL, L**
- What is constant? **temperature**



- Give an example to help explain the law (you should have some type of visual of your example).
- What would a graph of pressure vs. volume of a gas look like according to Boyle's Law?

Charles' Law: Answer the following questions about Charles' Law.

- Describe Charles' Law in your own words. **Increase temperature, increase volume**
- What is the formula? **$T_1/V_1 = T_2/V_2$**
- What are the variables? What are the possible units used in the formula? **Temperature and volume (kelvin, mL, L)**
- What is constant? **Pressure**
- What is very important to remember about the temperature variable when using this formula? **It must be in kelvin**
- Give an example to help explain Charles's Law (you should have some type of visual of your example).



- What would a graph of temperature vs. volume of a gas look like according to Charles's Law?

Gay-Lussac's Law: Answer the following questions about Gay-Lussac's Law.

- Describe Gay-Lussac's Law in your own words. **As pressure increases, temperature increases**
- What is the formula? **$P_1/T_1 = P_2/T_2$**
- What are the variables? What are the possible units used in the formula? **Pressure and temperature (kelvin, atm, kPa, torr)**
- What is constant? **volume**
- What is very important to remember about the temperature variable when using this formula? **It must be in kelvin**
- Give an example to help explain Gay-Lussac's Law (you should have some type of visual of your example).

Gas Laws Test Study Guide

James E. Brady, John R. Holm

Gas Laws Test Study Guide:

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Study Guide with Selected Solutions Karen Timberlake,2002-08-06

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Fundamentals of Chemistry, Study Guide James E. Brady,John R. Holum,1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics Emphasizes descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises

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General Chemistry, Study Guide James E.

Brady,1990-03-01 The Fifth Edition retains the pedagogical strengths that made the previous editions so popular and has been updated reorganized and streamlined Changes include more accessible introductory chapters with greater stress on the logic of the periodic table earlier introduction of redox reactions greater emphasis on the concept of energy a new section on Lewis structures earlier introduction of the ideal gas law and a new development of thermodynamics Each chapter ends with review questions and problems

General Chemistry Study Guide Sixth Edition Darrell D. Ebbing,1999

Study Guide [to Accompany] General Chemistry James E. Brady,1982

Study Guide Joan I. Senyk,Darrell D. Ebbing,1984

Study Guide for General Chemistry and College Chemistry, Ejighth Editions by Holtzclaw and Robinson Norman E. Griswold,Henry Fuller Holtzclaw,1988

Study Guide and Full Solutions Manual Susan McMurry,John

McMurry,2002-08 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all end of chapter problems in the textbook

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Paul B. Kelter,1993

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Study Guide and Partial Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry Susan McMurry,1996 This internationally acclaimed detective series is just the thing for lovers of thoseNumber One Ladieslooking for a darker more realistic view of Botswana Sue Baker Publishing News

Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e

Kenneth W. Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond's text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry **Study Guide and Solutions Manual for Fundamentals of General, Organic, and Biological Chemistry** Susan McMurry, 1992

Study Guide, Chemical Principles, Fifth Edition, Zumdahl Paul B. Kelter, Steven S. Zumdahl, 2004-04

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