

# Electronic Configuration

1. Use noble gases to write the abbreviated electron configurations of the following?

i. Li \_\_\_\_\_

ii. B \_\_\_\_\_

iii. O \_\_\_\_\_

iv. Mg \_\_\_\_\_

v. Cl \_\_\_\_\_

vi. V \_\_\_\_\_

vii. Se<sup>2-</sup> \_\_\_\_\_

viii. Cr<sup>2+</sup> \_\_\_\_\_

ix. I<sup>-</sup> \_\_\_\_\_

x. Ag \_\_\_\_\_

2. Identify the element with the following electron configuration

i. 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>2</sup> 3p<sup>6</sup> 4s<sup>2</sup> 3d<sup>10</sup> 4p<sup>4</sup> \_\_\_\_\_

ii. 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>2</sup> 3p<sup>6</sup> 4s<sup>2</sup> 3d<sup>10</sup> 4p<sup>6</sup> 5s<sup>2</sup> \_\_\_\_\_

iii. 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>2</sup> 3p<sup>6</sup> 4s<sup>2</sup> 3d<sup>10</sup> 4p<sup>6</sup> 5s<sup>1</sup> 4d<sup>7</sup> \_\_\_\_\_

iv. 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>2</sup> 3p<sup>6</sup> 4s<sup>2</sup> 3d<sup>10</sup> 4p<sup>6</sup> 5s<sup>2</sup> 4d<sup>10</sup> 5p<sup>6</sup> 6s<sup>2</sup> 4f<sup>14</sup> 5d<sup>10</sup> \_\_\_\_\_

v. 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>2</sup> 3p<sup>6</sup> 4s<sup>2</sup> 3d<sup>10</sup> 4p<sup>6</sup> 5s<sup>2</sup> 4d<sup>10</sup> 5p<sup>6</sup> 6s<sup>2</sup> 4f<sup>14</sup> 5d<sup>10</sup> 6p<sup>6</sup> 7s<sup>1</sup> \_\_\_\_\_

3. Determine if the following electron configurations are correct:

i. 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>2</sup> 3p<sup>6</sup> 4s<sup>2</sup> 4d<sup>10</sup> 4p<sup>6</sup> 5s<sup>1</sup> \_\_\_\_\_

ii. 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>3</sup> \_\_\_\_\_

iii. [Rn] 7s<sup>2</sup> 5f<sup>6</sup> 6d<sup>2</sup> \_\_\_\_\_

iv. [Ar] 5s<sup>2</sup> 4d<sup>10</sup> 5p<sup>3</sup> \_\_\_\_\_

v. [Xe] 6s<sup>2</sup> 4f<sup>10</sup> \_\_\_\_\_

4. Identify the following elements

i. 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>2</sup> \_\_\_\_\_

ii. 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>0</sup> \_\_\_\_\_

iii. [Ar] 4s<sup>2</sup> 3d<sup>10</sup> 4p<sup>5</sup> \_\_\_\_\_

iv. [Kr] 5s<sup>2</sup> 4d<sup>1</sup> \_\_\_\_\_

v. [Ne] 3s<sup>2</sup> 3p<sup>1</sup> \_\_\_\_\_

# Electron Configuration Worksheet And Lots More

## Answer Sheet

**G Orfield**

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