

## Lab: Single and Double Replacement Reactions

**Purpose:** To predict products and observe various single and double replacement reactions.

**Hypotheses:** Single Replacement Reactions: Use the **activity series** (p.268 - Honors) and your understanding of single replacement reactions to predict products for the following reactions. If no reaction will occur, then write "no rxn" on the product side of the equation. Remember to **BALANCE** the equation afterwards.

- 1)  $\text{HCl} + \text{Mg} \rightarrow$  \_\_\_\_\_
- 2)  $\text{HCl} + \text{Cu} \rightarrow$  \_\_\_\_\_ (hint: copper could be +2 in this compound)
- 3)  $\text{HCl} + \text{Fe} \rightarrow$  \_\_\_\_\_ (hint: iron will be +3 in this compound)
- 4)  $\text{HCl} + \text{Zn} \rightarrow$  \_\_\_\_\_

**Hypotheses:** Double Replacement Reactions: Use the **solubility chart** and your understanding of double replacement reactions to predict products for the following reactions (see Table 2 on the back of this paper...we're only fully writing out equations for 4 of the 16 you will test). If both of the possible products are aqueous, then write "no rxn" on the product side of the equation. Otherwise, clearly state which product will be aqueous and which will form a solid precipitate. Remember to **BALANCE** the equation afterwards.

- 1)  $\text{Na}_2\text{CO}_3(\text{aq}) + \text{Cu}(\text{NO}_3)_2(\text{aq}) \rightarrow$  \_\_\_\_\_
- 6)  $\text{NaOH}(\text{aq}) + \text{Fe}(\text{NO}_3)_3(\text{aq}) \rightarrow$  \_\_\_\_\_
- 11)  $\text{Na}_3\text{PO}_4(\text{aq}) + \text{AgNO}_3(\text{aq}) \rightarrow$  \_\_\_\_\_
- 16)  $\text{MgSO}_4(\text{aq}) + \text{Zn}(\text{NO}_3)_2(\text{aq}) \rightarrow$  \_\_\_\_\_

### Procedures:

1. For each single replacement reaction, place a sample of each metal in your well plate and then place 5-6 drops of HCl (aq) on top of it. Record your observations.
2. For each double replacement reaction, place two drops of each solution carefully on the transparency, covering each "X" appropriately over your "Super High-Tech Patented Reaction Surface." Record your observations.
3. If a precipitate formed, clearly write the formula for the precipitate on the line provided.
4. Clean up your workstation according to your instructor's directions and answer all questions.

### Observations/Data:

Table 1: Single Replacement Reaction Observations

|    | HCl (aq)        |
|----|-----------------|
| Mg | 1 Observations: |
| Cu | 2 Observations: |
| Fe | 3 Observations: |
| Zn | 4 Observations: |

# Lab 27 Double Replacement Reactions

**Brian Pressley**



## Lab 27 Double Replacement Reactions:

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