

EXPERIMENT 9: Double Replacement Reactions

PURPOSE

- To identify the ions present in various aqueous solutions.
- To systematically combine solutions and identify the reactions that form precipitates and gases.
- To become familiar with writing equations for reactions, including net ionic equations.

EQUIPMENT & CHEMICALS

Spot plates (or small test tubes)

Stirring rod or toothpicks

0.1 M solutions of the following:

AgNO_3

$\text{Ca}(\text{NO}_3)_2$

CuSO_4

$\text{Fe}(\text{NO}_3)_3$

$\text{HC}_2\text{H}_3\text{O}_2$

HCl

K_2CrO_4

KI

NaCl

$\text{NaC}_2\text{H}_3\text{O}_2$

Na_2CO_3

NaOH

Na_2S

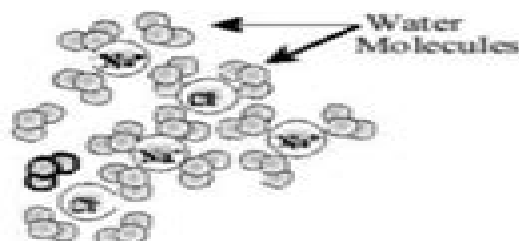
NH_4Cl

$\text{Pb}(\text{NO}_3)_2$

INTRODUCTION

When one substance dissolves in another substance, a **solution** is formed. A **solution** is a homogeneous mixture in which the components are uniformly mixed. A solution consists of **solute** (the species that is dissolved) and **solvent** (the medium in which the solute has dissolved). The solvent is usually present in larger amount than the solute. When water is the solvent, the solution is called an **aqueous solution**.

When an ionic compound dissolves in water, it dissociates into its constituent ions. Such a compound is a strong electrolyte and conducts electricity well in dilute aqueous solutions. For example, when NaCl dissolves in water, it dissociates into separate Na^+ and Cl^- ions.



This process occurs as polar water molecules orient themselves around the sodium and chloride ions and pull them free from the solid crystal. Once removed from the solid crystal, the ions remain separated and surrounded by water molecules. There are no solid NaCl particles present, although in the solution, occasional contact between the Na^+ and Cl^- ions, called **ion pairing**, does occur.

The ability of a substance to dissolve is called its **solubility**. The solubility of a substance in water is typically defined as the mass of the substance that will dissolve in 100 mL of water. Solubility is temperature-dependent and usually increases for solid solutes dissolving in a liquid solvent.

Lab 9 2 Double Replacement Reactions

Colleen F. Craig, Kim N. Gunnerson

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the authors' names.

Lab 9 2 Double Replacement Reactions:

CHEMISTRY EXPERIMENTS James Signorelli, 2014-09-19 Gifted and talented students and any student interested in pursuing a science major in college needs a rigorous program to prepare them while they are still in high school This book utilizes a format where the application of several disciplines science math and language arts principles are mandated Each lab concludes with either an essay or a detailed analysis of what happened and why it happened This format is based on the expectations of joining a university program or becoming an industrial science professional The ideal student lab report would be written in a lab research notebook and then the essay or final analysis is done on a word processor to allow for repeat editing and corrections The research notebook has all graph pages a title section and a place for the students and their assistants to sign and witness that exercise The basic mechanics of the lab report title purpose procedure diagrams data table math and calculations observations and graphs are handwritten into the book The conclusion is done on a word processor MS Word which allows the instructor to guide the student in writing and editing a complete essay using the MLA format When the final copy is completed the essay is printed and inserted into the lab notebook for grading At the end of the term the student has all their labs in one place for future reference These lab notebooks can be obtained for as little as 3 00 per book This is money well spent In our district the Board of Education buys the books for each student The BOE sees these books as expendable but necessary materials for all science and engineering instruction Merrill Chemistry-Lab.Manual Smoot, 1994-07

Foundations of College Chemistry, Laboratory Morris Hein, Susan Arena, 2010-08-09 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis **Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 Prentice Hall Chemistry, 2000

Exploring General Chemistry in the Laboratory Colleen F. Craig, Kim N. Gunnerson, 2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science **Laboratory Experiments for General, Organic & Biochemistry** Frederick A. Bettelheim, Joseph Marvin Landesberg, 1997 Laboratory Manual for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new

edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh,2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application *Foundations of Chemistry in the Laboratory* Morris Hein,Leo R. Best,Robert L. Miner,1973 **Chemistry** Eugene LeMay, Jr.,Herbert Beall,Karen M. Robblee,Douglas C. Brower,Prentice-Hall Staff,2002-02 Elementary principles of chemistry Raymond Bedell Brownlee,1921 **Chemistry Workbook and Laboratory Guide** Martin Van Buren McGill,1933 **Subject Guide to Books in Print** ,2001 **Laboratory Equipment for Recognized High Schools in Illinois ...** Illinois. Office of the Superintendent of Public Instruction,1926 **Laboratory Manual for Fundamentals of Chemistry 3/E** Jo A. Beran,1988 Merrill Laboratory Chemistry L. Neal Carmichael,David F. Haines,1990-10 *Chemistry* Karen Timberlake,1999 Suitable for one or two term lab courses covering general organic and biological chemistry this new edition written by Karen Timberlake features many improvements to the insightful experiments that have made it the leading lab manual Each experiment encourages critical thinking with laboratory goals discussion of related concepts clear instructions new pre lab questions and comprehensive report pages Forty one experiments illustrate the basic principles of chemistry **A Laboratory Manual of Organic Chemistry for Beginners** Arnold Frederick Holleman,1913 **Chemistry Guide and Laboratory Exercises** Martin Van Buren McGill,G. M. Bradbury,1935

The book delves into Lab 9 2 Double Replacement Reactions. Lab 9 2 Double Replacement Reactions is a crucial topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Lab 9 2 Double Replacement Reactions, encompassing both the fundamentals and more intricate discussions.

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 - Chapter 3: Lab 9 2 Double Replacement Reactions in Everyday Life
 - Chapter 4: Lab 9 2 Double Replacement Reactions in Specific Contexts
 - Chapter 5: Conclusion
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3. In chapter 2, this book will delve into the foundational concepts of Lab 9 2 Double Replacement Reactions. This chapter will elucidate the essential principles that must be understood to grasp Lab 9 2 Double Replacement Reactions in its entirety.
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6. In chapter 5, the author will draw a conclusion about Lab 9 2 Double Replacement Reactions. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Lab 9 2 Double Replacement Reactions.

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