

# Experiment 34 Prelaboratory Assignment

## An Equilibrium Constant

Date \_\_\_\_\_ Lab Sec. \_\_\_\_\_ Name \_\_\_\_\_ Desk No. \_\_\_\_\_

1. Three parameters affect the absorbance of a sample. Which one is the focus of this experiment?

(1) Concentration of the solution

(2) Tube length

(3) Wavelength

Concentration is being focused on

2. Experimental Procedure, Part A.1, Table 34.1. A 3.00-ml. aliquot of 0.001 M NaSCN is diluted to 25.0 ml. with 0.2 M  $\text{Fe}(\text{NO}_3)_3$  and 0.1 M  $\text{HNO}_3$ .

a. How many moles of  $\text{SCN}^-$  are present?

$$\text{no. of moles of } \text{SCN}^- = \frac{0.001 \text{ M} \times 3 \text{ ml}}{1000 \text{ ml}} = 3 \times 10^{-6} \text{ mol}$$

- b. If all of the  $\text{SCN}^-$  is complexed with  $\text{Fe}^{3+}$  to form  $\text{FeNCS}^{2+}$ , what is the molar concentration of  $\text{FeNCS}^{2+}$ ?

$$3 \times 10^{-6} \text{ mol of } \text{SCN}^- \text{ produce } 3 \times 10^{-6} \text{ mol of } \text{FeNCS}^{2+}$$

Total volume of diluted solution = 25 ml = 0.025 L

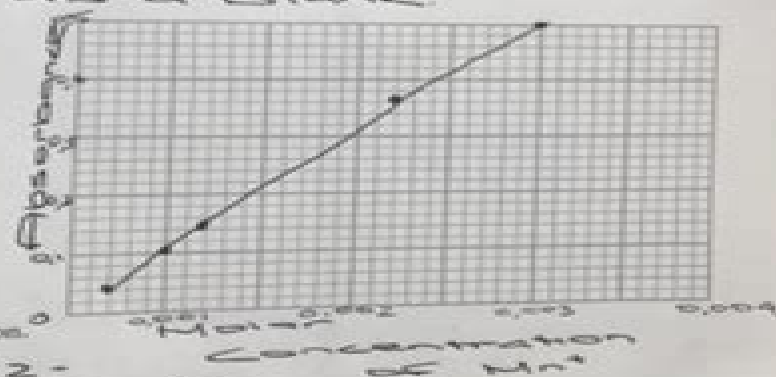
3. Experimental Procedure, Part A.1. For preparing a set of standard solutions of  $\text{FeNCS}^{2+}$ , the equilibrium molar concentration of  $\text{FeNCS}^{2+}$  is assumed to equal the initial molar concentration of the  $\text{SCN}^-$  in the reaction mixture. Why is this assumption valid? The  $\text{SCN}^-$  is the limiting reagent and thus when all of  $\text{SCN}^-$  is added its consumed no more complex formed.

4. Experimental Procedure, Part A.3. The blank solution used to calibrate the spectrophotometer is 10.0 ml. of 0.2 M  $\text{Fe}(\text{NO}_3)_3$  diluted to 25.0 ml. with 0.1 M  $\text{HNO}_3$ . Why is this solution preferred to simply using de-ionized water for the calibration? The blank prepared above also quenchs absorbance which must be subtracted to the correct value for the complex.

When the actual reaction is done, so, blank water cannot be used as a blank.

5. Plot the following data as absorbance versus  $[\text{M}^{3+}]$  as a calibration curve:

Absorbance, A	Molar Concentration of $\text{M}^{3+}$
0.045	$3.0 \times 10^{-4} \text{ mol/L}$
0.097	$6.2 \times 10^{-4} \text{ mol/L}$
0.14	$9.0 \times 10^{-4} \text{ mol/L}$
0.35	$2.2 \times 10^{-3} \text{ mol/L}$
0.51	$3.2 \times 10^{-3} \text{ mol/L}$



A test solution showed a percent transmittance (%T) reading of 38.4%T. Interpret the calibration curve to determine the molar concentration of  $\text{M}^{3+}$  in the test solution.

$$\begin{aligned} \text{Absorbance} &= 2 - 10\log\%T = 2 - 10\log(38.4) = 0.416 \\ \text{Concentration of unknown solution} &= 0.416 \times 3.2 \times 10^{-3} / 0.51 = 2.61 \times 10^{-3} \text{ mol/L} \end{aligned}$$

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## Answers

**M Lipman**



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## **Table of Contents Experiment 34 Equilibrium Constant Report Sheet Answers**

1. Understanding the eBook Experiment 34 Equilibrium Constant Report Sheet Answers
  - The Rise of Digital Reading Experiment 34 Equilibrium Constant Report Sheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Experiment 34 Equilibrium Constant Report Sheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Experiment 34 Equilibrium Constant Report Sheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Experiment 34 Equilibrium Constant Report Sheet Answers

- Personalized Recommendations
  - Experiment 34 Equilibrium Constant Report Sheet Answers User Reviews and Ratings
  - Experiment 34 Equilibrium Constant Report Sheet Answers and Bestseller Lists
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    - Experiment 34 Equilibrium Constant Report Sheet Answers Budget-Friendly Options
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    - ePub, PDF, MOBI, and More
    - Experiment 34 Equilibrium Constant Report Sheet Answers Compatibility with Devices
    - Experiment 34 Equilibrium Constant Report Sheet Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Experiment 34 Equilibrium Constant Report Sheet Answers
    - Highlighting and Note-Taking Experiment 34 Equilibrium Constant Report Sheet Answers
    - Interactive Elements Experiment 34 Equilibrium Constant Report Sheet Answers
  8. Staying Engaged with Experiment 34 Equilibrium Constant Report Sheet Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Experiment 34 Equilibrium Constant Report Sheet Answers
  9. Balancing eBooks and Physical Books Experiment 34 Equilibrium Constant Report Sheet Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Experiment 34 Equilibrium Constant Report Sheet Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Experiment 34 Equilibrium Constant Report Sheet Answers
    - Setting Reading Goals Experiment 34 Equilibrium Constant Report Sheet Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Experiment 34 Equilibrium Constant Report Sheet Answers

- Fact-Checking eBook Content of Experiment 34 Equilibrium Constant Report Sheet Answers
- Distinguishing Credible Sources

### 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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