

Experiment 34 Prelaboratory Assignment

An Equilibrium Constant

Date _____ Lab Sec. _____ Name _____

Desk No. _____

1. Three parameters affect the absorbance of a sample. Which one is the focus of this experiment?

(1) Concentration of the solution
 (2) Tube length
 (3) Wavelength

2. Experimental Procedure, Part A.1. Table 34.1. A 1.00-mL aliquot of 0.001 M NaSCN is diluted to 25.0 mL with 0.2 M $\text{Fe}(\text{NO}_3)_3$, and 0.1 M HNO₃.

a. How many moles of SCN⁻ are present?

$$\text{no. of moles of SCN}^- = \frac{0.001}{1000} \times 25.0$$

$$= 2.5 \times 10^{-5} \text{ moles}$$

b. If all of the SCN⁻ is complexed with Fe^{3+} to form FeNCS^{2+} , what is the molar concentration of FeNCS^{2+} ?
 $2.5 \times 10^{-5} \text{ mol of SCN}^- \text{ produce } 2.5 \times 10^{-5} \text{ mol of FeNCS}^{2+}$

Total volume of diluted solution = 25.0 mL

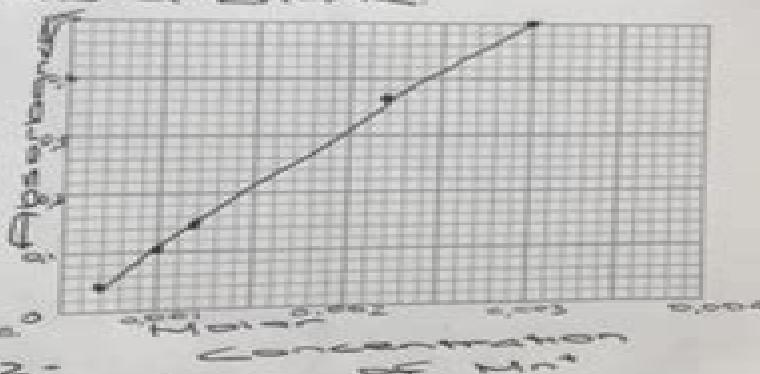
Concentration of FeNCS^{2+} = $\frac{2.5 \times 10^{-5}}{25.0} = 2.5 \times 10^{-6} \text{ M}$

3. Experimental Procedure, Part A.1. For preparing a set of standard solutions of FeNCS^{2+} , the equilibrium molar concentration of FeNCS^{2+} is assumed to equal the initial molar concentration of the SCN⁻ in the reaction mixture. Why is this assumption valid? The SCN⁻ is the limiting reagent and thus when all of SCN⁻ is reacted, there is no more SCN⁻ to react with Fe^{3+} . The concentration of SCN⁻ remains the same as the concentration of FeNCS^{2+} .

4. Experimental Procedure, Part A.3. The blank solution used to calibrate the spectrophotometer is 10.0 mL of 0.2 M $\text{Fe}(\text{NO}_3)_3$, diluted to 25.0 mL with 0.1 M HNO₃. Why is this solution preferred to simply using de-ionized water for the calibration? The ~~de-ionized~~ blank prepared above also gives a reference value which must be subtracted to the correct value for the samples from when the actual reaction is done. So, plan how samples will be used as a blank.

5. Plot the following data as absorbance versus $[\text{M}^{3+}]$ as a calibration curve.

Absorbance, A	Molar Concentration of M^{3+}
0.045	$3.0 \times 10^{-5} \text{ mol/L}$
0.097	$6.2 \times 10^{-5} \text{ mol/L}$
0.14	$9.0 \times 10^{-5} \text{ mol/L}$
0.25	$2.2 \times 10^{-4} \text{ mol/L}$
0.51	$3.2 \times 10^{-3} \text{ mol/L}$



A test solution showed a percent transmittance (%T) reading of 38.4% T. Interpret the calibration curve to determine the molar concentration of M^{3+} in the test solution.

$$\text{Concentration} = 2 \cdot 10^{-7} \times T^{-2} - \frac{1}{10^7 (38.4) \times 0.416}$$

Concentration of unknown solution Experiment 34 383

$$\text{Concentration of unknown} = 0.416 \times 3.2 \times 10^{-7} / 0.384 - 2.4 \times 10^{-7} \text{ mol/L}$$

Experiment 34 Equilibrium Constant Report Sheet

Answers

M Lipman

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1. Understanding the eBook Experiment 34 Equilibrium Constant Report Sheet Answers
 - The Rise of Digital Reading Experiment 34 Equilibrium Constant Report Sheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Experiment 34 Equilibrium Constant Report Sheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Experiment 34 Equilibrium Constant Report Sheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Experiment 34 Equilibrium Constant Report Sheet Answers

- Personalized Recommendations
- Experiment 34 Equilibrium Constant Report Sheet Answers User Reviews and Ratings
- Experiment 34 Equilibrium Constant Report Sheet Answers and Bestseller Lists

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- Highlighting and Note-Taking Experiment 34 Equilibrium Constant Report Sheet Answers
- Interactive Elements Experiment 34 Equilibrium Constant Report Sheet Answers

8. Staying Engaged with Experiment 34 Equilibrium Constant Report Sheet Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Experiment 34 Equilibrium Constant Report Sheet Answers

9. Balancing eBooks and Physical Books Experiment 34 Equilibrium Constant Report Sheet Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Experiment 34 Equilibrium Constant Report Sheet Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Experiment 34 Equilibrium Constant Report Sheet Answers

- Setting Reading Goals Experiment 34 Equilibrium Constant Report Sheet Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Experiment 34 Equilibrium Constant Report Sheet Answers

- Fact-Checking eBook Content of Experiment 34 Equilibrium Constant Report Sheet Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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web dec 15 1999 flora ii in persia with approximately six thousand recorded species of ferns and flowering plants persia harbors one of the richest floras of the near eastern countries which is surprising given that more than two thirds of the country's surface consists of deserts semideserts and steppes

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web the geographic area of flora iranica covers the entire political boundaries of iran and afghanistan n iraq mountains of turkmenistan small parts of azerbaijan talesh and the western