

4) Oxidation-Reduction Reactions

Oxidation - is a loss of electrons
 - is an increase in oxidation number

Reduction - is a gain of electrons
 - is a decrease in oxidation number

Oxidizing Agent - brings about the oxidation of another species
 ↗ gets reduced

Reducing Agent - brings about the reduction of another species
 ↗ causes something else to gain electrons by losing
 ↗ gets oxidized

→ OIL RIG (Oxidation Is Loss; Reduction Is Gain)
 LEO~GER (Lose Electrons Oxidation ~ Gain Electrons Reduction)

Memorize common oxidizing pathways (reducing goes backward)

1. Metal $\xrightarrow{\text{ox}}$ metal ion
 $\text{Zn} \rightarrow \text{Zn}^{2+}$

2. Metal $\xrightarrow{\text{ox}}$ metalous ion $\xrightarrow{\text{ox}}$ metallic ion
 $\text{Cu} \rightarrow \text{Cu}^{+} \rightarrow \text{Cu}^{2+}$

3. Halide ion $\xrightarrow{\text{ox}}$ halogen $\xrightarrow{\text{ox}}$ halite ion $\xrightarrow{\text{ox}}$ halate ion (similar for $\text{NO}_2 \rightarrow \text{NO}_3^-$)
 $\text{Cl}^- \rightarrow \text{Cl}_2 \rightarrow \text{ClO}^- \rightarrow \text{ClO}_2^- \rightarrow \text{ClO}_3^- \rightarrow \text{ClO}_4^-$

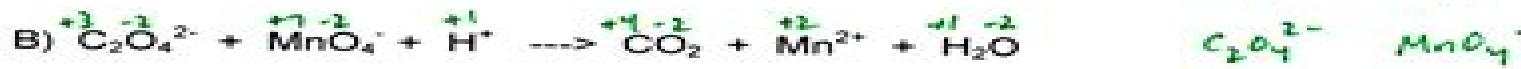
Memorize common oxidizing / reducing agents (see bright orange sheet)

Ex. $\text{KMnO}_4 \rightarrow \text{Mn}^{2+}$ (acidified solution)

$\text{MnO}_4^- \rightarrow \text{MnO}_2$ (basic solution)

$\text{Cr}_2\text{O}_7^{2-} \rightarrow \text{Cr}^{3+}$ (acidified solution)

==== Identify the substances being oxidized and reduced in each of the following. ===



Experiment 46 Oxidation Reduction Reactions Answers

George Curtis Martin

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AP Chemistry For Dummies Peter J. Mikulecky, Michelle Rose Gilman, Kate Brutlag, 2008-11-13 A practical and hands on guide for learning the practical science of AP chemistry and preparing for the AP chem exam Gearing up for the AP Chemistry exam AP Chemistry For Dummies is packed with all the resources and help you need to do your very best Focused on the chemistry concepts and problems the College Board wants you to know this AP Chemistry study guide gives you winning test taking tips multiple choice strategies and topic guidelines as well as great advice on optimizing your study time and hitting the top of your game on test day This user friendly guide helps you prepare without perspiration by developing a pre test plan organizing your study time and getting the most out of your AP course You'll get help understanding atomic structure and bonding grasping atomic geometry understanding how colliding particles produce states and so much more To provide students with hands on experience AP chemistry courses include extensive labwork as part of the standard curriculum This is why the book dedicates a chapter to providing a brief review of common laboratory equipment and techniques and another to a complete survey of recommended AP chemistry experiments Two full length practice exams help you build your confidence get comfortable with test formats identify your strengths and weaknesses and focus your studies You'll discover how to Create and follow a pretest plan Understand everything you must know about the exam Develop a multiple choice strategy Figure out displacement combustion and acid base reactions Get familiar with stoichiometry Describe patterns and predict properties Get a handle on organic chemistry nomenclature Know your way around laboratory concepts tasks equipment and safety Analyze laboratory data Use practice exams to maximize your score Additionally you'll have a chance to brush up on the math skills that will help you on the exam learn the critical types of chemistry problems and become familiar with the annoying exceptions to chemistry rules Get your own copy of AP Chemistry For Dummies to build your confidence and test taking know how so you can ace that exam

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